WHAT IS CLAIMED IS:

1. A non-aqueous electrolytic solution comprising glycerine and at least one soluble salt formed by the neutralization of at least one non-halogen-containing organic or inorganic acid anion with at least one alkali metal, ammonium, or protonated amine cation; wherein the acid anion is derived from an acid having a pKa lower than phosphoric acid.

2. The non-aqueous electrolytic solution according to claim 1 wherein the soluble salt is ammonium nitrate, dimethyl ethanolamine nitrate, dimethyl ethanolamine sulfate, dimethylethoxy ethanolamine nitrate, or dimethylethoxy ethanolamine sulfate.

3. The non-aqueous electrolytic solution according to claim 2 wherein the soluble salt is ammonium nitrate.

4. The non-aqueous electrolytic solution according to claim 1 wherein water content is less than 2 wt%, based on total weight of the solution.

5. The non-aqueous electrolytic solution according to claim 4 wherein water content is less than 1 wt%, based on total weight of the solution.

6. The non-aqueous electrolytic solution according to claim 1 comprising about 0.5 wt% to about 15 wt% of the soluble salt, based on total weight of the solution.

7. The non-aqueous electrolytic solution according to claim 6 comprising about 5 wt% to about 10 wt% of the soluble salt , based on total weight of the solution.

8. A non-aqueous electrolytic solution comprising glycerine and ammonium nitrate.

9. A method of anodizing an anode comprising anodizing at a temperature of about 60°C to about 125°C until a uniform anodic oxide film is formed over the entire anode

surface with the non-aqueous electrolytic solution according to claim 1; wherein the anode comprises a valve metal-derived nitride, sub-nitride, oxide, or sub-oxide, or an alloy thereof, a mixture thereof, or a metallic glass composition thereof.

10. The method according to claim 9 wherein the temperature is about 80°C to about 95°C.

11. The method according to claim 10 wherein the temperature is about 84°C to about 92°C.

12. The method according to claim 9 wherein the anode comprises tantalum nitride, niobium nitride, or titanium nitride.

13. A capacitor comprising an anode prepared from a valve-metal derivative powder and a non-aqueous electrolytic solution comprising glycerine and at least one soluble salt formed by the neutralization of at least one non-halogen-containing organic or inorganic acid anion with at least one alkali metal, ammonium, or protonated amine cation; wherein the acid anion is derived from an acid having a pKa lower than phosphoric acid, and wherein the valve-metal derivative powder is a valve metal-derived nitride, sub-nitride, oxide, or sub-oxide, or an alloy thereof, a mixture thereof, or a metallic glass composition thereof.

14. The capacitor according to claim 13 wherein the soluble salt is ammonium nitrate, dimethyl ethanolamine nitrate, dimethyl ethanolamine sulfate, dimethylethoxy ethanolamine nitrate, or dimethylethoxy ethanolamine sulfate.

15. The capacitor according to claim 14 wherein the soluble salt is ammonium nitrate.

16. The capacitor according to claim 13 wherein water content of the solution is less than 2 wt%, based on total weight of the solution.

17. The capacitor according to claim 16 wherein water content of the solution is less than 1 wt%, based on total weight of the solution.

18. The capacitor according to claim 13 wherein the solution comprises about 0.5 wt% to about 15 wt% of the soluble salt, based on total weight of the solution.

19. The capacitor according to claim 18 wherein the solution comprises about 5 wt% to about 10 wt% of the soluble salt, based on total weight of the solution.

20. The capacitor according to claim 13 wherein the valve-metal derivative is tantalum nitride, niobium nitride, or titanium nitride.